1. **Use the Table of Ion Charges to name the following compounds:**

**eg. PbCO3  : Pb = Lead, CO3 = Carbonate ... PbCO3 is Lead Carbonate**

1. KNO3
2. FeSO4
3. NH4OH
4. Fe2(SO4)3
5. HgCl2
6. MgSO4
7. **How many atoms of oxygen are there in:**
8. 1 molecule of SO2
9. 3 molecules of H2SO4
10. 5 "molecules" of Fe3(PO4)2
11. 100 molecules of CH3COOH
12. 10 "molecules" of Ca(HCO3)2
13. **List the number and each kind of atom in each of:**
14. 3 molecules of ethanol, C2H5OH
15. 3 "molecules" of calcium phosphate, Ca3(PO4)2
16. 4 "molecules" of chromium (III) sulphate, Cr2(SO4)3
17. 2 "molecules" of magnesium chlorate, Mg(ClO3)2
18. **Use the table of Ion Charges on the previous page to write the chemical formulas for:**

**eg. Zinc Bromide:**

**Zinc Zn 2+, Bromide Br 1-**

**ZnBr2**

1. Potassium sulphate
2. Calcium carbonate
3. Magnesium oxide
4. Copper (I) chloride
5. Ammonium hydroxide
6. Barium nitrite
7. Silver chloride
8. Sodium nitrate
9. **Copy out the following table and fill in the blanks using your periodic table:**

|  |  |  |  |
| --- | --- | --- | --- |
| **Atomic Number** | **Element** | **Symbol** | **Possible Ion it forms** |
| **4** | **Beryllium** | **Be** | **Be2+** |
| **13** |  |  |  |
| **35** |  |  |  |
| **37** |  |  |  |
|  |  | **Ba** |  |
|  |  |  | **At -** |
|  | **Strontium** |  |  |

1. **Complete the crossword below.**

|  |
| --- |
| usestitle |
|  | **Across**  **5** He (6) **6** Ca (7) **8** O (6) **9** Fe (4) **10** Cu (6) **12** K (9) **15** Na (6) **16** Au (4) **18** H (8) **19** Be (9) **24** Cl (8) **25** S (7) **26** Ne (4) |
| **Down**  **1** N (8) **2** Ni (6) **3** Al (9) **4** As (7) **7** Zn (4) **11** P (10)1. Pu (9)

**13** Ar (5) |  **14** Si (7) **17** F (8)1. B (5)

 **20**  Pb (4) **21** Hg (7)1. Ag (6)

**23** C (6) | atommm |
|  |  |  |  |  |